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## Practical Medicinal Chemistry Laboratory Manual

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MEDICINAL CHEMISTRY IIA INTRODUCTION The practical part of the course covers experiments on the synthesis of organic compounds and drugs using different analysis and purification techniques and includes synthesis of chiral drugs and the techniques used in the separation of enantiomers. EQUIPMENT AND INSTRUMENTS Analytical Balances

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PROCEDURE: STEP-1: In a 2 lit RBF fitted with a double surface reflux condenser place, 11.5gm (0.5gm atom) of clean sodium and it is dissolved in 250ml of absolute alcohol. To this solution add 80gm (0.5mole of ethyl malonate) followed by 30gm (0.5mole of dry urea) dissolved in 250ml of ethanol.

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The units of volume commonly used in this manual are  $\text{cm}^3$  and  $\text{dm}^3$ . The conversion between  $\text{cm}^3$  and  $\text{dm}^3$  is  $1 \text{ dm}^3 = 1000 \text{ cm}^3$ . Note that a  $\text{cm}^3$  are the same as millilitres (ml) and  $\text{dm}^3$  are the same as litres (L).  $1 \text{ cm}^3 = 1 \text{ ml}$  and  $1 \text{ dm}^3 = 1 \text{ L}$ . B. Concentration Concentration is a measure of amount per volume. The amount can be grams, kilograms,

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Medicinal Chemistry Laboratory Manual by Charles Dickson

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1) Collect the following items which you will need to carry out the experiment. (a) 2 100mL round bottom flask (b) 1 spatula (c) 1 cork ring (d) 1 plastic bowl (e) 1 ceramic Buchner funnel (f) 1 watch glass (g) 1 magnetic stirrer bar (h) 1 zip lock bag. 2) Put on a pair of gloves before you start the experiment.

PR1110 Lab Manual - Foundations for Medicinal Chemistry ...

Step 1: First step involves the formation of Acetylum ion, the reaction is enhanced by proton donation by concentrated acid. Step 2: Amino group of Para-Aminophenol dissociates as follows: Step 3: Acetylum ion attacks the anion at the nitrogen atom to form Para-acetyl Aminophenol or Paracetamol.

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