

Chapter 8 Chemistry Covalent Bonding

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1. In covalent bonds, electron sharing usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively charged. The attractive forces between the two atoms result from the

Chapter 8: Covalent Bonding and Molecular Structure

1) Ionic bonds—transfer of electrons. (gained or lost; makes formula unit) 2) Covalent bonds—sharing of electrons. The resulting particle is called a “molecule”. Covalent Bonds. •The word covalent is a combination of the prefix co- (from Latin com, meaning “with” or “together”), and the verb. valere, meaning “to be strong”.

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A covalent bond in which the bonding electrons are most likely to be found in sausage-shaped regions above and below the bond axis of the bonded atoms. Polar molecule. A molecule in which one side of the molecule is slightly negative and the opposite side is slightly positive. Polar covalent bond.

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CHAPTER 8 SOLUTIONS MANUAL Covalent Bonding Covalent Bonding Solutions Manual
Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH_3 2. H_2S 3. HCl respectively, for single, double, and triple bonds.

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Complete Table 8.10, which shows the number of electrons shared in a single covalent bond, a double covalent bond, and a triple covalent bond. Identify the group of atoms that will form each of these bonds.

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View full document. Chapter 8 Covalent Bonding. The Covalent Bond Atoms will share electrons in order to form a stable octet. Covalent bond: the chemical bond that results from the sharing of valence electrons also called a molecular bond. Molecule: formed when two or more atoms bond covalently. Formed when the proton of one atom is attracted to the electron cloud of another atom.

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

Pearson Chemistry Chapter 8 Covalent Bonding Answers

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Chapter 8 Covalent Bonds 1. Covalent Bonding Chapter 8 2. The atoms held together by sharing electrons are joined by a Covalent Bond. Sharing is Caring 3. 2. Covalent bonds- Two atoms share one or more pairs of outer-shell electrons. Oxygen Atom Oxygen Atom Oxygen Molecule (O_2) 4.

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