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~~autoionization of H_2O~~

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16 Acids Bases

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Identifying Bronsted

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~~and 18. Acid/base~~

~~and OX-RED~~

~~reactions. Acid Base~~

~~Relationships~~

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~~16 (Acid-Base Equilibria) – Part 3~~
Chapter 16 Review
Acid Base

Section 16.1 □ ACIDS AND BASES: A BRIEF REVIEW □

Acids and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red, whereas bases turn

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litmus blue. Acid

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Chapter 16: Acid

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BASES. Table 16.1.

1: General Properties
of Acids and Bases.

produce a piercing
pain in a wound. give
a slippery feel. taste
sour. taste bitter. are
colorless when placed
in phenolphthalein (an

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indicator). are pink when placed in phenolphthalein (an indicator). are red on blue litmus paper (a pH indicator).

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Chapter 16. 16.1 Acids and Bases: A Brief Review; 16.2

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Bronsted-Lowry Acids and Bases; 16.3 The Autoionization of Water; 16.4 The pH Scale; 16.5 Strong Acids and Bases; 16.6 Weak Acids; 16.7 Weak Bases; 16.8 Relationship Between K_a and K_b ; 16.9 Acid-Base Properties of Salt Solutions; 16.10 Acid-Base Behavior and

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Chemical Structure;

16.11 ...

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Answers

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Section 1

Chapter 16: Acid and

Base Review

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~PLEASE DO NOT
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What two substances are always produced by a neutralization reaction? a. acid and a base b. water and a base c. water and an acid d. water and a salt 2.

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Acid and Base

Review

Chapter 16 Acids and

Bases 1. Acids were

recognized primarily
from their sour taste.

Bases were
recognized from their
bitter taste and
slippery feel on skin.

2. In the Arrhenius
definition, an acid is a
substance that
produces hydrogen

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ions (H^+) when dissolved in water, whereas a base is a substance that produces hydroxide ions (OH^-) in

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16 Acids and Bases

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Bases: A Brief Review

16.2 Brønsted-Lowry

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Acids and Bases 16.3

The Autoionization of
Water 16.4 The pH

Scale 16.5 Strong

Acids and Bases 16.6

Weak Acids 16.7

Weak Bases 16.8

Relationship between

K_a and K_b 16.9 Acid-

Base Properties of

Salt Solutions 16.10

Acid-Base Behavior

and Chemical

Structure 16.11 Lewis

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Acids and Bases Ch.

16 Mastering

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Overview The

Arrhenius definition is

the narrowest view of

...

Chapter 16 - Acid

Base Equilibrium -

Chapter 16 Acid-Base

...

16.1 Acids and

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Bases: A Brief Review

□ Acids taste sour and cause certain dyes to change color. □ Bases taste bitter and feel soapy. □ Arrhenius concept of acids and bases: □ An acid is a substance that, when dissolved in water, increases the concentration of H^+ ions. □ Example: HCl is an acid. □ An

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Arrhenius base is a substance that, when dissolved in water, increases the concentration of OH^- ions.

AP Chemistry

CHAPTER 16 STUDY

GUIDE Acid-Base

Equilibrium

Acids and Bases Acid

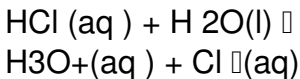
and Base Strength In

any acid-base

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reaction, the equilibrium will favor the reaction that moves the proton to the stronger base.



H_2O is a much stronger base than Cl^- , so the equilibrium lies so far to the right K is not measured ($K \gg 1$).

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Chapter 16 Acids and
Bases

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16.10: Acid-Base

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Behavior and

Chemical Structure

Inductive effects and charge delocalization

significantly influence the acidity or basicity

of a compound. The

acid-base strength of a molecule depends

strongly on its

structure. The weaker

the $A-H$ or $B-H^+$

bond, the more likely

it is to dissociate to

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form an (H^+) ion.

Base Ratio Ph

16: Acid-Base

Equilibria - Chemistry

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CHAPTER 16 □ Acid-

Base Equilibria

Section 16.1 □ Acids

and Bases: A Brief

Review (a) Define an

acid and a base,

according to the

Arrhenius definition.

acid = base = (b)

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Write the products of each chemical reaction below, which involves the dissociation of each reactant into aqueous ions.

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Acid-Base Equilibria

...

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AND BASES: A

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This video explains

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Review the concepts from
your packet on
Chapter 16 (Acid-
Base Equilibria),
which can be found
here:

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Section 16.1: Acids
an...

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Equilibria - YouTube
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16.1 Acids & Bases: A Brief Review. □

Arrhenius acids and
bases: □□ acid: an H^+
donor $HA \rightleftharpoons H^+ + A^-(aq)$

$(aq) \rightleftharpoons (aq) + OH^-(aq)$ □□ base: an
 OH^- donor $MOH \rightleftharpoons M^+ + OH^-(aq)$

Brønsted-Lowry acids
and bases: □□ acid: an
 H^+ donor $HA \rightleftharpoons H^+ + A^-(aq)$
 $(aq) \rightleftharpoons (aq) + OH^-(aq)$ □□ base: an
 H^+ acceptor HB

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everyday applications, the authors explain chemical concepts by starting with the basics, using symbols or diagrams, and conclude by encouraging

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This general

chemistry video

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tutorial focuses on acids and bases and buffer solutions. It shows you how to calculate the pH and pOH of the solution. It cont...

Ka Kb Kw pH pOH
pKa pKb H⁺ OH⁻
Calculations - Acids ...

This Chapter 15

Review, Section 2:

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and pH Worksheet is suitable for 9th - 12th Grade. Keep it simple with this chemistry assignment. Learners examine an acid-base titration graph and answer four questions about the data.

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